

Disproportionation.

One and the same substance may act simultaneously as an oxidizing agent and as a reducing agent with the result that a part of it gets oxidized to a higher state and rest of it is reduced to lower state of oxidation. Such a reaction, in which a substance undergoes simultaneous oxidation and reduction is called **disproportionation** and the substance is said to **disproportionate**.

Following are the some examples of disproportionation,

