The Screening Effect or Shielding Effect.

A valence-electron in a multi-electron atom is attracted by the nucleus, and repelled by the electrons of inner-shells. The combined effect of this attractive and repulsive force acting on the valence-electron experiences less attraction from the nucleus. This is called **shielding** or **screening effect**. The magnitude of the screening effect depends upon the number of inner electrons, i.e., higher the number of inner electrons, greater shall be the value of screening effect. The screening effect constant is represented by the symbol ' σ ' is determined by the **Slater's rules**.

The magnitude of screening constant in the case of s- and p- block elements increases in a period as well as in a group as the atomic number increases.